

# Study Guide The Mole Answers

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## Study Guide The Mole Answers

TEACHER GUIDE AND ANSWERS Chemistry: Matter and Change Teacher Guide and Answers 8 molar mass  $KCO_3$  99.11 g/mol  $KCO_3$  n molar mass of molecular formula/molar mass of empirical formula 198.22 g/mol/99.11 g/mol  $2(KCO_3)_n$  The molecular formula of the compound is  $K_2C_2O_6$ . Section 10.5 The Formula for a Hydrate 1. hydrate 2. hydration

## Ch 10 Study Guide TE

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The Mole Unit In laboratory practice, you work with much larger quantities of the elements than single atoms or single molecules. A convenient standard quantity is the mole, the amount of the substance equal in grams to the sum of the atomic masses. Therefore, one mole of carbon dioxide is 44.01 grams:

### **The Mole Unit - CliffsNotes Study Guides**

Know how to use the molecular mass to find the number of moles of a compound, given the mass in grams and vice-a-versa converting empirical formula to a mol(g→Mol) Know how to use the mass percent to find the empirical formula 40.0% C, 6.7% H, 53.5% O. Molecular mass of 60.0g/mol.

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In your textbook, read about chemical formulas and the mole, the molar mass of compounds, and conversions among mass, moles, and number of particles. Study the table and the diagram of a methane molecule and a trichloromethane molecule. Then answer the following questions.

### **VIBRATIONS AND WAVES**

A mole of a particular substance is equal to the number of atoms in exactly 12 g of the carbon-12 isotope. Experiments established that number to be  $6.022142 \times 10^{23}$  particles. Avogadro's number =  $N_A = 6.022 \times 10^{23}$  items/mole. In other words, one mole equals  $6.022 \times 10^{23}$  items. The units of the mole are modified depending on the units needed in a calculation.

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Calculate If a mole is  $6.02 \times 10^{23}$  items, how far will a mole of your items, placed end-to-end length-wise, extend into space? Express your answer in meters. 2. Calculate Convert the distance in Question 1 to light-years (ly). (1 ly =  $9.46 \times 10^{15}$  m) 3. Compare the distance you calculated in Question 2 with these astronomical distances: a.

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Mass of each element or compound 1) Sn 11878 g/mole 2) Fe 2 O 3 15939 g/mole 3) H 2 SO 4 9808 g/mole Calculate the mass in grams 4) 8 moles of NH 3 136 g 8 mol NH3 x 17 g NH3 5) 10 moles of AgNO 3 16987 g [EPUB] Chapter 11 Study Guide Answers The Mole

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